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lnK = (?rS/R) - (?rH/RT) The equilibrium for the dissociation of one mole of iodine vapour is 3.94 x10⁻⁴ at 900K and the standard enthalpy is 154 kJmol⁻¹. Calculate the equilibrium constant at 1200K assuming the standard enthalpy remains constant. ln (K2/K1) = - ?rH/R (1/T2 - 1/T1) , ANS = 6.7 x 10⁻² at 1200K.

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